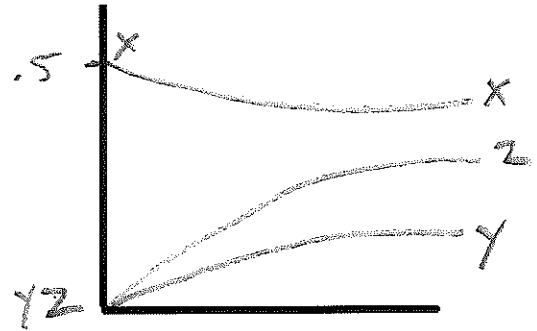
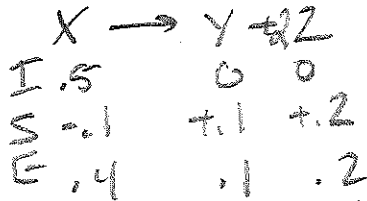


Introduction to Completion vs. Equilibrium

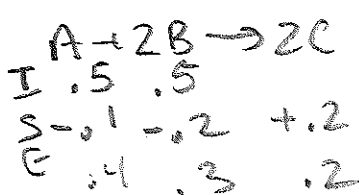
1. A Chemical process of 0.5M of X only goes 20% to product producing Y and Z. Write a balanced reaction with an ISC table showing ISE stoichiometry.



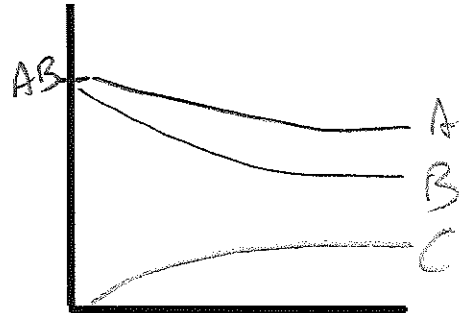
2. Draw a graphical representation of this process?

A chemical process of $A + 2B \leftrightarrow 2C$. A and B both start at a concentration of 0.5M and when the process has come to equilibrium 20% of A is converted.

3. Write the equilibrium expression for this reaction.
 4. Determine K_c for this process.
 5. Draw a graphical representation of concentration.

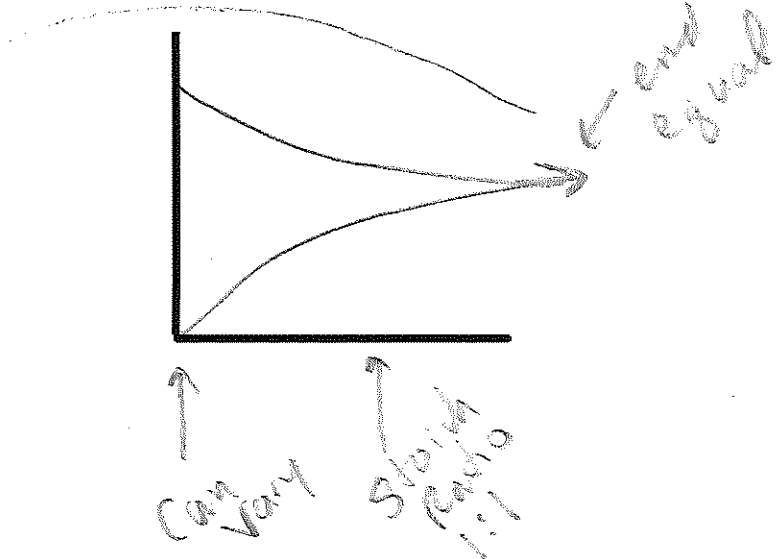


$$K_c = \frac{[C]^2}{[A][B]^2}$$



- 6.

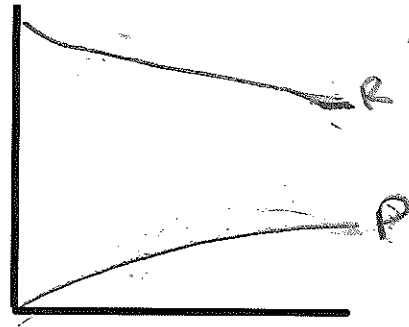
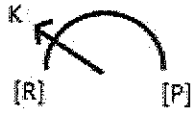
STP $[R] = 1M$ $[P] = 1M$



7.



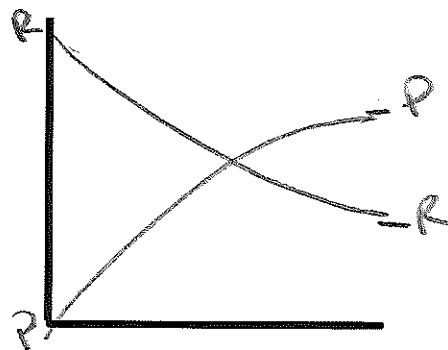
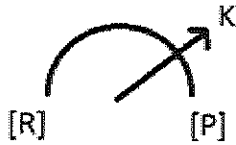
Reactant Favored Reaction



8.



Product Favored Reaction



9. Using the graph draw an arrow on the gauge indicating placement of equilibrium ratios.

